

## CHIKMAGALUR FOREST DIVISION.

Notification dated 7th July 1952.

**No. 1-52-53.** It is hereby notified for the general information of the public, that the undermentioned timber was found adrift, near 'Kanthi' on Chikmagalur-Koppa Road and the same has been taken possession by the Forest Department.

No.	Kind	Length Ft. in.	Girth Ft. in.	Cft.
1	Bite	13 6	6 8	33
2	"	16 0	4 4	20
3	"	13 3	4 6	17
4	"	9 0	3 10	8
5	"	14 0	4 6	16
6	"	12 3	3 6	9
7	"	11 0	6 0	25

Any person having a claim over the above timber, will prefer a written statement with evidence to the District Forest Officer, Chikmagalur Division, Chikmagalur, within one month from the date of this Notification, failing which, the timber will be confiscated to Government.

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K. M. SUBBAYYA,  
District Forest Officer.

## CHITALDRUG FOREST DIVISION.

Notification dated 15th July 1952.

It is hereby notified for the general information of the public that the undermentioned kind and quantity of timber found unclaimed in Lakkihalli State Forest, is collected and kept under the custody of the Forest Department.

## Kamara Timber.

- A. (1)  $6'-0'' \times 2'-4'' = 2-0-6$   
 (2)  $15'-0'' \times 1'-9'' = 2-10-5$   
 (3)  $8'-0'' \times 2'-4'' = 2-8-8$   
 (4)  $7'-0'' \times 2'-7'' = 2-11-0$   
 (5)  $7'-0'' \times 2'-9'' = 3-5-1$

Total ... 5 Pieces = 13-11-3 or 14 c.ft.

## Kamara Timber.

- B. (1)  $5'-3'' \times 2'-4'' = 1-9-5$   
 (2)  $7'-6'' \times 2'-3'' = 2-4-4$   
 (3)  $8'-0'' \times 2'-6'' = 3-1-6$   
 (4)  $8'-0'' \times 2'-5'' = 2-11-0$   
 (5)  $8'-0'' \times 2'-5'' = 2-11-0$   
 (6)  $11'-6'' \times 2'-1'' = 3-1-6$   
 (7)  $11'-0'' \times 2'-0'' = 2-9-0$   
 (8)  $11'-0'' \times 2'-3'' = 3-5-9$   
 (9)  $8'-0'' \times 2'-5'' = 2-11-0$

Total ... 9 Pieces = 25-4-6 or 25 c.ft.

The claimants, if any, for the above timber, are required to apply to the District Forest Officer, Chitaldrug, within a period of one month from the date of this Notification, for considering their claims.

If no claims are preferred within the time allowed above, the timber will be transferred to Government Account and the claims, if any, preferred thereafter will be rejected.

C. R. A. PADMANABHA RAO,  
D.F.O.

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## UNIVERSITY OF MYSORE.

Notification No. Ex. 2-267-51-52 of 3rd August 1951.

## Amendment No. 77 to Calendar Vol. I for 1947-48.

The following are the revised titles of papers II and III in History for the B.A. Degree Examination:—  
 Page 193 of Calendar Vol. I for 1947-48:

## 11 History.

For

- (ii) History of India from 1300 to 1920.  
 (iii) History of Europe from 1500 to 1920.

Substitute:

- (i) History of India from 1300.  
 (iii) History of Europe from 1500.

(Letter No. E. 3710, dated 19th July 1951, from the Education Secretary to Government, Bangalore.)

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[Pages 175 to 177 of University Calendar, Vol. I for 1947-48.]

Notification No. Ex. 3-764/51-52, dated 8th February 1952.

## Syllabus in Chemistry for the Intermediate Examination in Science.

## PART I. (To be completed in the Junior Intermediate Class).

Elements, compounds, mixtures; physical and chemical changes: types of chemical change: Law of conservation of energy, Law of conservation of mass, Laws of chemical combination by weight.

Equivalent weights of elements, methods of determining equivalent weights, Dalton's atomic theory, explanation of the laws of chemical combination according to the atomic theory.

Boyle's law, Charles's law the gas equation, Dalton's law of partial pressures, Graham's law of diffusion. Gay Lussac's Law of combining volumes, Avogadro's hypothesis: relation between molecular weight and vapour density, the gram molecular volume, the method of finding molecular weights of gases [Regnault's method] and Volatile liquids and solids [Victor Meyer's method.]

Determination of atomic weights using Dulong and Petit's law, the Law of isomorphism, and the concept of the least weight in a molecule.

Symbols, formulae, equations: valency, radicals, calculation of reacting quantities: determination of empirical formulae from percentage composition: molecular and structural formulae.

Electrolytes and non-electrolytes; electrolysis, Faraday's Laws; elementary ideas of ionic theory.

Hydrogen.—Preparation of water, acids and alkalis; preparation of pure hydrogen, manufacture, properties, uses.

Oxygen.—Discovery of oxygen: the nature of combustion preparation of oxygen from the atmosphere and from compounds; properties and uses; classification of oxides.

Ozone.—Preparation, properties and uses; molecular formula.

Hydrogen Peroxide.—Preparation, molecular formula, properties, uses, estimation of hydrogen peroxide.

Water.—Natural waters; purification of water for different purposes; hardness of water; methods softening; the volumetric composition of steam; Dumas and Morley's experiments on the hydrogen-oxygen ratio; hydrates: efflorescence and deliquescence; drying agents.

Solubility of solids in liquids: solubility curves; solubility of gases in liquids; Henry's Law.

Reversible reactions; chemical equilibrium. Endothermic and Exothermic reactions.

Oxidation and reduction; acids, bases and salts; acid, basic and normal salts; neutralisation; indicators; equivalent weights of compounds; standard solutions.

Halogens.—Preparation and properties of the halogen elements and their hydrides; manufacture and uses of chlorine and hydrogen chloride; composition hydrogen chloride, comparison of the halogens.

Sulphur.—Extraction, allotropy, properties, uses.

Hydrogen Sulphide.—Preparation, properties, composition, uses.

Sulphur Dioxide.—Preparation, properties, composition, uses, sulphurous acid.

Sulphur Trioxide.—Preparation and properties.

Sulphuric Acid.—Manufacture, properties and uses.

Nitrogen.—The composition of the atmosphere; preparation of nitrogen from the air and from compounds; properties and uses.

Ammonia.—Preparation, properties, uses, composition; Le Chatelier's principle; manufacture of Ammonia by Haber's process; ammonium salts:—Chloride, sulphate, carbonate and nitrate; thermal dissociation.

Oxides of Nitrogen.—Preparation and properties; composition of nitrous and nitric oxides.

Nitric Acid.—Preparation, properties, uses; manufacture by Ostwald's process, nitrates.

Nitrous Acid.—Nitrites.

Phosphorus.—Allotropy, preparation, properties and uses of white and red phosphorus; preparation and properties

phosphine, chlorides of phosphorus, phosphorus trioxide, phosphorus pentoxide, phosphorus acid, the phosphoric acids, phosphorus oxychloride.

**Arsenic.**—Preparation and properties of arsenious oxide, arsine, arsenic trichloride.

## PART II (To be completed in the Senior Intermediate Class).

Elementary ideas of radioactivity, the structure of the atom, isotopes, the periodic classification of elements.

**Boron.**—Preparation, properties and uses of boric acid and borax.

**Silicon.**—Silica, preparation and properties of silicon tetrafluoride and sodium silicate; glass.

Colloidal solutions: Dialysis, coagulation, Silica gel.

**Carbon.**—Allotropy, uses of active charcoal; preparation and properties of carbon monoxide, fuels, composition of carbon dioxide, carbon disulphide.

The occurrence, properties and uses of the following metals and their compounds. The preparation of the compounds to be indicated.

**Sodium.**—Metal, oxide, peroxide, hydroxide, carbonate, bicarbonate, halides, hypochlorite, sulphite, thiosulphate, cyanide, phosphate, nitrate.

Manufacture of common salt; the hydroxide, bicarbonate and carbonate of sodium.

**Potassium.**—Nitrate, chlorate, perchlorate.

**Copper.**—Metal, electrolytic refining; alloys, oxides, chlorides, copper sulphate.

**Silver.**—Metal, oxide, nitrate, halides.

**Gold.**—Metal and gold chloride.

**Magnesium.**—Metal, chloride, oxide, sulphate.

**Calcium.**—Metal, oxide, hydroxide, carbonate, chloride, sulphate, carbide, cyanamide, phosphate, superphosphate, bleaching powder (manufacture).

**Barium.**—Oxide, peroxide, hydroxide, carbonate, chloride, sulphate.

**Zinc.**—Metal, oxide, chloride, sulphate.

**Mercury.**—Metal, oxides, chlorides, sulphate, nitrates.

**Aluminium.**—Metal, oxide, chlorides, sulphate, alum.

**Tin.**—Metal, oxides, chlorides.

**Lead.**—Metal, oxides, chlorides, nitrate, acetate, white lead, alloys of lead.

**Chromium.**—Chromic oxide, chromium trioxide, chromates, dichromates, chrome alum.

**Manganese.**—Dioxide, potassium permanganate.

**Iron.**—Cast iron, wrought iron, steel; oxides, ferrous and ferric chlorides, ferrous sulphate, ferrous ammonium sulphate, potassium ferrocyanide, potassium ferricyanide, double and complex salts.

## ORGANIC CHEMISTRY.

**Introduction.**—Characteristics of organic compounds.

**Methane, ethane, Homologous series common petroleum products ethylene, acetylene, methyl iodide, chloroform, carbon tetrachloride, methyl alcohol, ethyl alcohol, glycerol, ether, formaldehyde, acetaldehyde, glucose, cane sugar, formic acid, acetic acid, acetone, ethyl acetate, benzene, toluene, naphthalene, phenol benzaldehyde, benzoic acid, salicylic acid, nitrobenzene, aniline.**

**Elementary account of starch, cellulose, oils and fats, proteins (egg albumin and casein).**

## PRACTICAL CHEMISTRY (TWO YEARS' COURSE).

1. Cutting, bending and drawing out glass tubing, study of Bunsen flame and the Burner.

2. Preparation and study of the properties of oxygen, oxides, hydrogen peroxide, hydrogen, chlorine, hydrogen chloride, ammonia, nitrous and nitric oxides, nitric acid, hydrogen sulphide, sulphur dioxide, carbon dioxide, boric acid, alum, ferrous ammonium sulphate, potassium permanganate and sodium thiosulphate.

3. Balance, sensitivity of the balance, Determination of water of hydration in crystals of zinc sulphate and barium chloride.

Equivalent weight of a metal by displacement of hydrogen or another metal; and by combination with oxygen or chlorine.

Solubility of a salt in water at 30°C or 35°C; solubility curves.

Molecular weights of oxygen and carbon dioxide.

4. Oxidation and reduction reactions.

Reactions of sodium hydroxide, sodium carbonate and ammonium hydroxide with metallic salts.

Study of the properties of hard and soft waters.

5. Qualitative analysis of simple salts. (Soluble in water or dil. acids.)

6. Volumetric analysis: exercises involving the use of the following standard solutions:—

Sodium carbonate, oxalic acid, hydrochloric acid, sodium hydroxide, permanganate and thiosulphate.

[Letter No. E. 10787, dated 8th February 1952, from the Educational Secretary to Government, Bangalore—Effective from the Junior Class of 1952-53.]

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Notification No. Ex. 2-1012-51-52, dated 18th June 1952.

## Amendment No. 115 to Calendar Vol. I for 1947-48.

The following is the detailed courses of Study for the Diploma courses in Persian and Arabic:—

### PERSIAN

**First Year:—**Elements of Grammar, and Pronunciation, the use of simple sentences, translation (seen and unseen passages) and reading of a text-book.

**Second Year:—**Grammar continued, more advanced translation, reading of prescribed text-books, conversation and free composition.

### TEXT BOOKS.

Extracts from prose works of Sadi, Jami, Hussain Waiz Kashifi from Safina-i-Adab.

Wukala-i-Murafa'ah, Romeo and Juliet by Ali Asghar Hikmat.

### BOOKS FOR REFERENCE.

Persian Grammar by Platts and Ranking.

Persian to English Dictionary by Steingas.

History of Persian Literature by Riza Shafag.

### ARABIC

**First Year:—**Elements of Grammar, and pronunciation, the use of simple sentences, translation (seen and unseen passages) and reading of a text-book.

**Second Year:—**Grammar continued, more advanced translation, reading of prescribed text-books, conversation, and free composition.

### TEXT BOOKS.

Al-Qiratur' Rashidah, part II and III.

Teach Yourself Arabic by Tritton.

Majani'ul Adab, Chapters I and II.

Tajir-i-Baghdad by Kamal Gilani.

### BOOKS FOR REFERENCE.

Arabic Literature by Gibb.

Arabic-English Dictionary by Hava.

Arabic Grammar by Thornton and Nicholson.

(Letter No. E. 3236, dated 18th June 1952, from the Education Secretary to Government, Bangalore.)

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Notification dated 23rd-24th July 1952.

**No. Ex. 5-392-52-53.** The following is the Special Subject prescribed in Psychology for the Honours Degree Final Examinations of 1954 and 1955.

SPEARMAN: "ABILITIES OF MAN".

By Order,

C. RANGACHAR,

Registrar.

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